



Evolution of hydrogeochemical properties during reactive flow-through experiments using CO₂-bearing solution

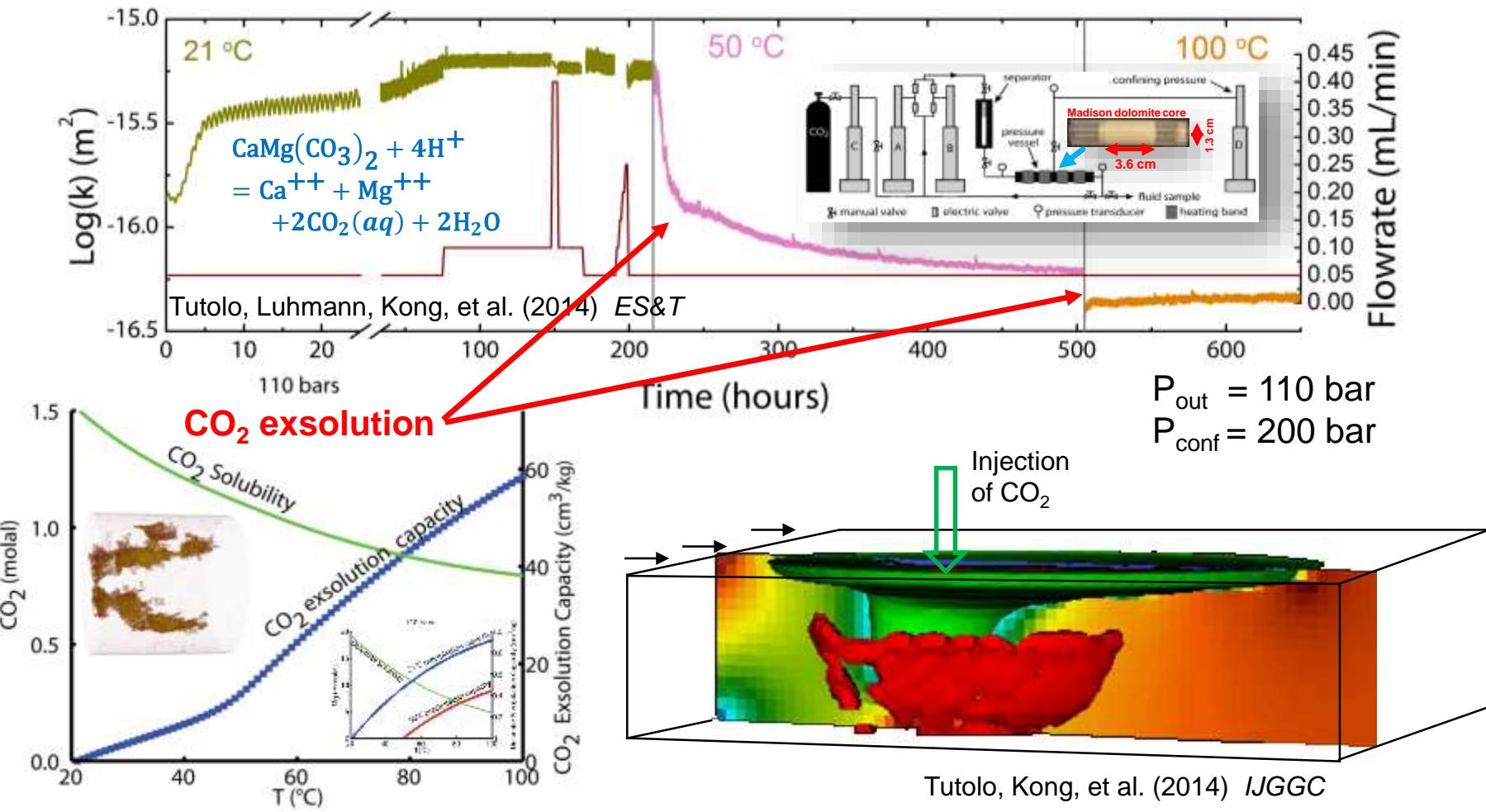
Xiang-Zhao Kong*, Jin Ma, Martin O. Saar

*Email: xkong@ethz.ch

Outline

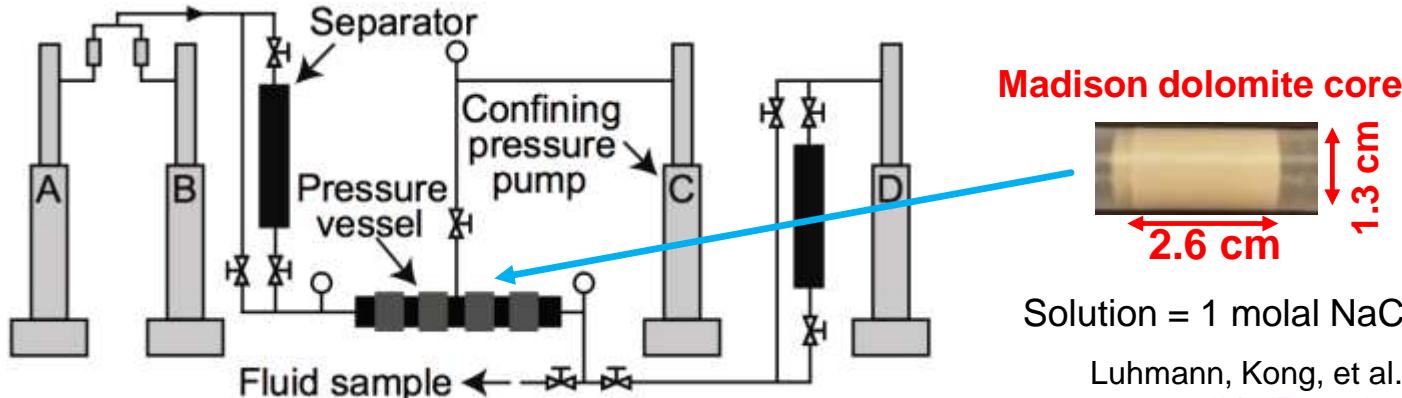
- **Introduction**
Previous work and motivation
- **Reactive flow-through experiments**
Experimental setup and results
- **Numerical simulation**
Simulation setup and results
- **Outlooks**

Precipitation due to temperature change



Dolomite dissolution

Single-pass experiments: far-from equilibrium

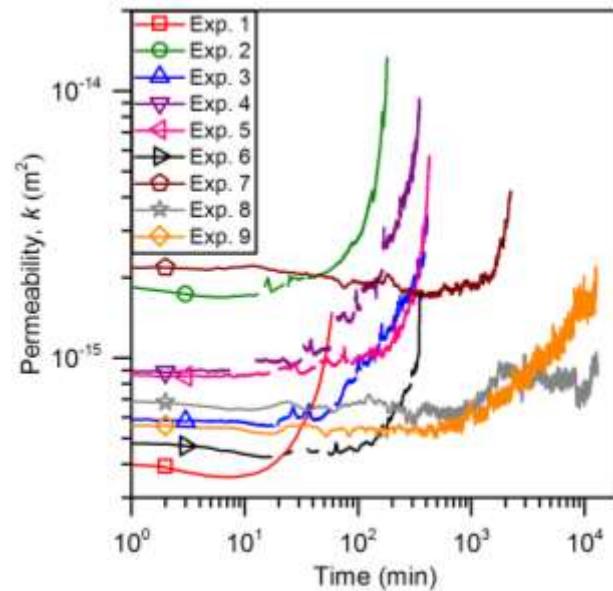


$$\begin{aligned} P_{\text{conf}} &= 200 \text{ bar} \\ P_{\text{out}} &= 150 \text{ bar} \\ T &= 100 \text{ }^{\circ}\text{C} \end{aligned}$$

Solution = 1 molal NaCl + 0.6 mol CO₂/kg

Luhmann, Kong, et al. (2014) Chem. Geol.

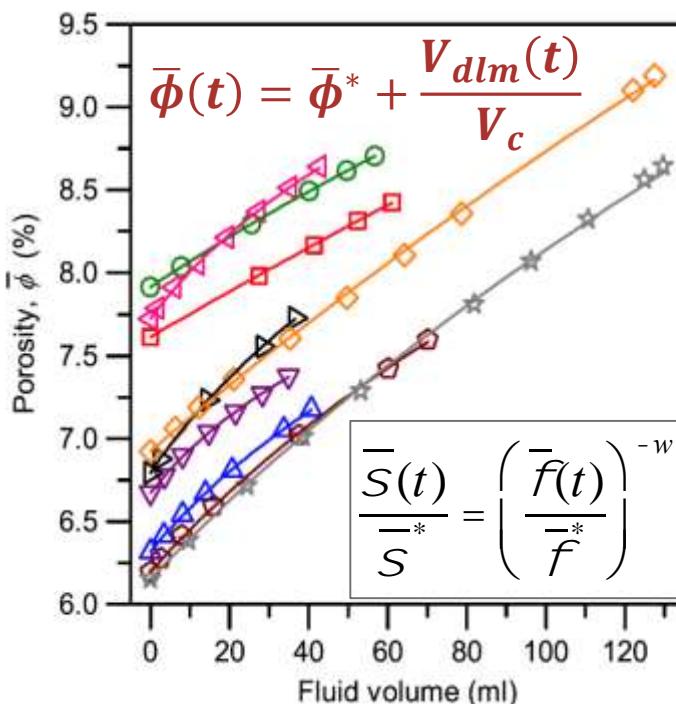
No.	Flowrate (ml/min)	Duration (min)	Pre-exp. mass (g)	Dissolved mass (g)	Initial permeability (m ²)	Final permeability (m ²)
1	1	61	8.213	0.146	3.7×10^{-16}	2.79×10^{-15}
2	0.316	180	7.921	0.114	1.31×10^{-15}	2.12×10^{-14}
3	0.1	408	8.037	0.112	5.7×10^{-16}	4.3×10^{-15}
4	0.1	350	8.012	0.094	7.7×10^{-16}	7.9×10^{-15}
5	0.1	427	8.001	0.118	7.5×10^{-16}	1.07×10^{-14}
6	0.1	367	7.833	0.12	4.5×10^{-16}	2.05×10^{-15}
7	0.0316	2,222	7.950	0.162	1.47×10^{-15}	3.8×10^{-15}
8	0.01	12,962	7.949	0.273	4.5×10^{-16}	1.32×10^{-15}
9	0.01	12,742	8.163	0.268	2.89×10^{-16}	1.32×10^{-15}



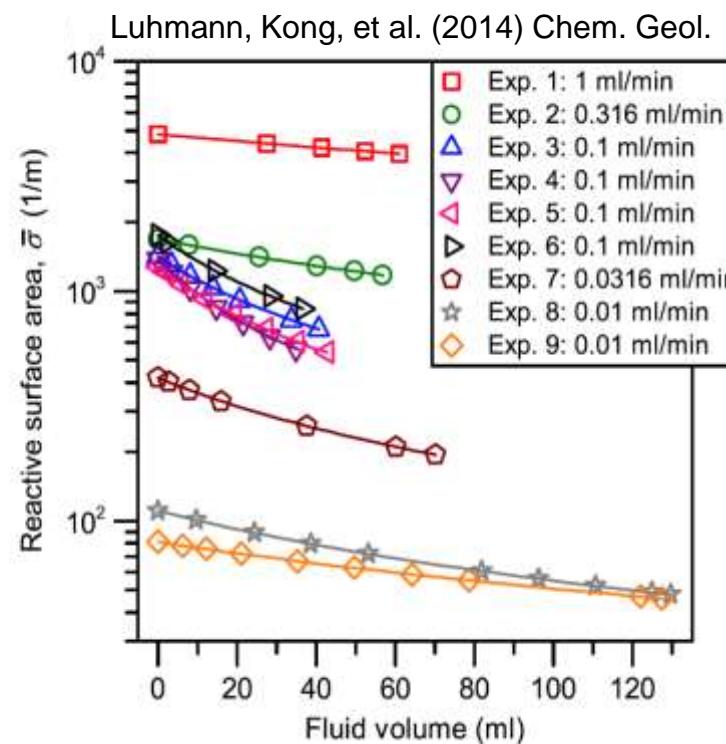
Dolomite dissolution

Single-pass experiments: surface area v.s. porosity

$$V_{dlm}(t) = 0.5vQ_v\rho_f \int_{\tau=0}^{\tau=t} [x\Delta C_{Ca}(\tau) + y\Delta C_{Mg}(\tau)] d\tau$$



Specific surface area determined by the BET method is at the order of 10^5 [1/m]



Initial porosity determined by X-ray Computed Tomography with a resolution of 8 um



U : dolomite molar volume

r_f : fluid density

ΔC_{Ca} : Ca concentration

ΔC_{Mg} : Mg concentration

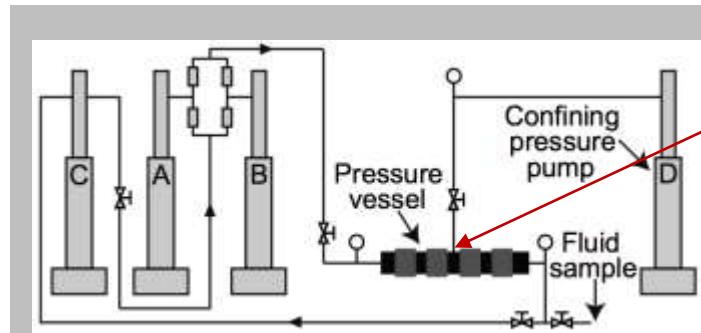
$$\begin{aligned} \bar{R}(t) &= \frac{1}{Uf(t)} \frac{\partial \bar{f}(t)}{\partial t} \\ &= \bar{S}(t) \left[1 - \frac{\bar{Q}}{K} \right] \end{aligned}$$

$$\begin{aligned} \bar{f}(t) &= \bar{f}^* [1 + tB(w)]^{1/w} \\ \bar{S}(t) &= \frac{\bar{S}^*}{1 + tB(w)} \end{aligned}$$

$$B(w) = wU/rS^* \left[1 - \frac{\bar{Q}}{K} \right]$$

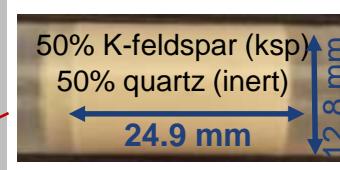
w: time-resolved dissolution efficiency.

Dissolution of K-feldspar-rich sandstone



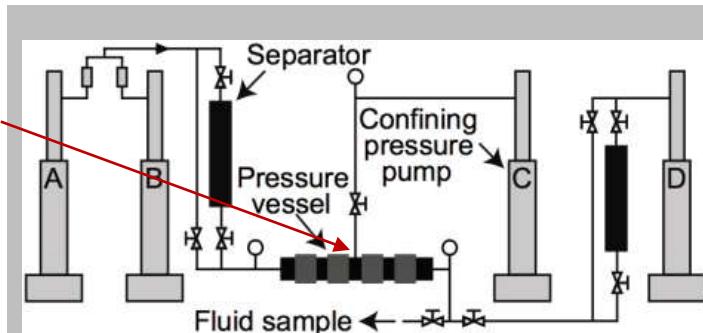
Recycling (RC) experiment:
approaching equilibrium

Deionized water + 0.86 mol CO₂/kg
Q=0.5 mL/min



$$\begin{aligned} P_{\text{out}} &= 200 \text{ bar} \\ P_{\text{conf}} &= 240 \text{ bar} \\ T &= 150 \text{ }^{\circ}\text{C} \end{aligned}$$

Tutolo, Luhmann, Kong, et al. (2015) GCA



Single-pass (SP) experiments:
far-from equilibrium

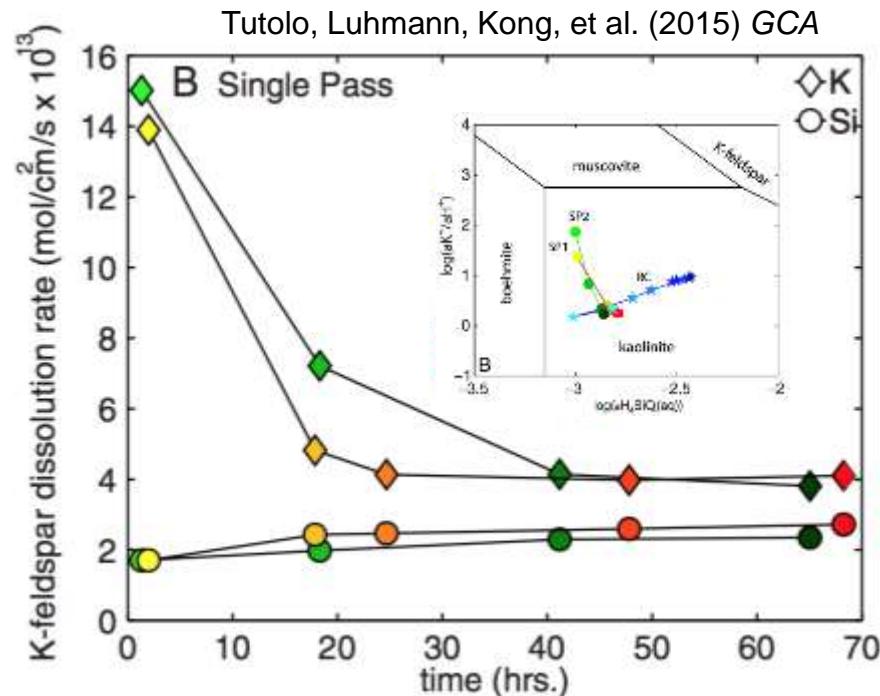
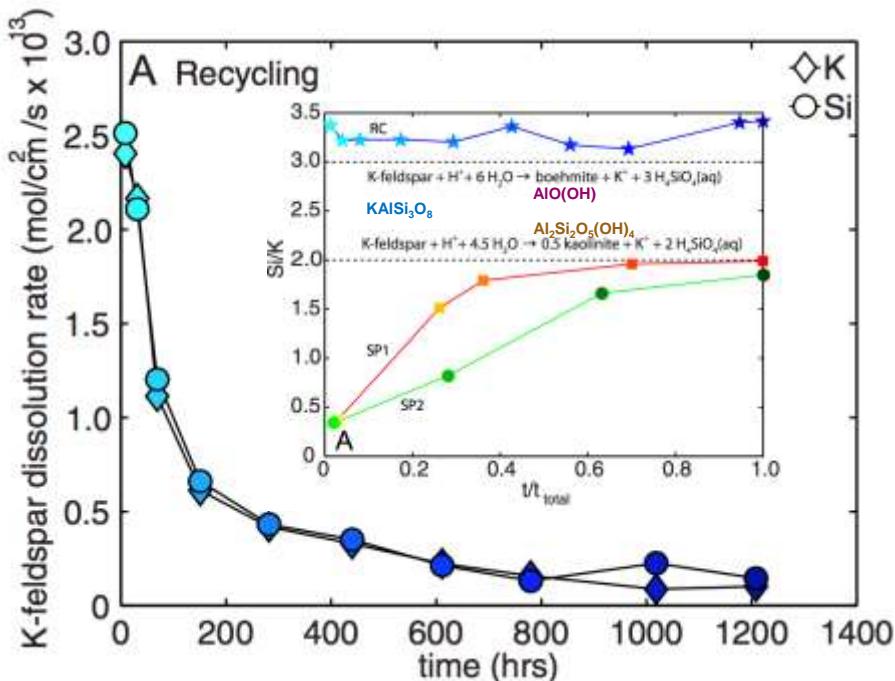
0.94 mol NaCl/kg + 0.75 mol CO₂/kg
Q=0.1 mL/min

Summary of parameters measured on the three cores before and after the experiments.

Experiment	Mass (g)			Permeability (10 ⁻¹⁴ m ²)		
	Pre	Post	Change (%)	Pre	Post	Change (%)
Recycling	5.632	5.509	-2.2	5.7	3.6	-37
Single-pass 1	-	-	-	2.3	1.6	-30
Single-pass 2	5.659	5.538	-2.1	4.7	3.5	-26

Reduction

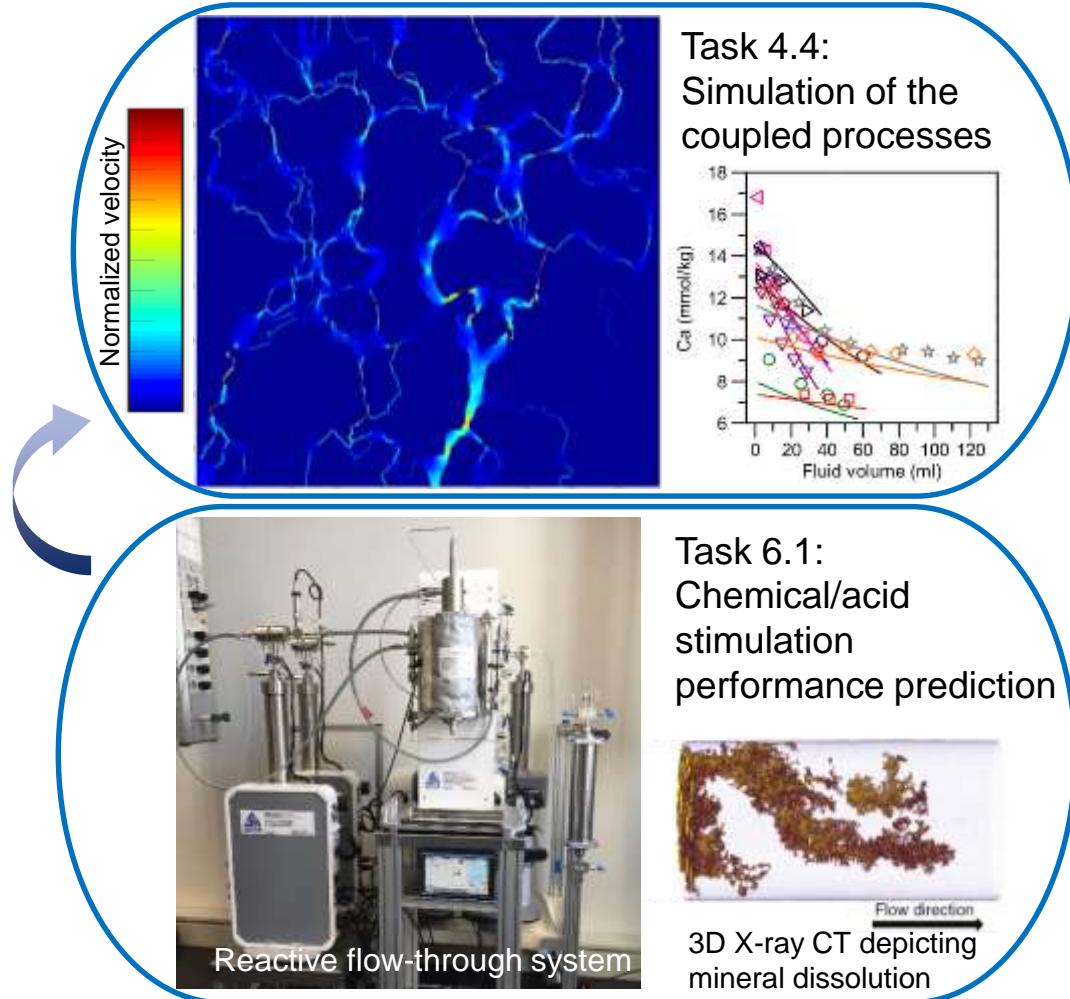
Dissolution of K-feldspar-rich sandstone



$$r = \frac{\Delta C Q}{\nu P S_{ksp}}$$

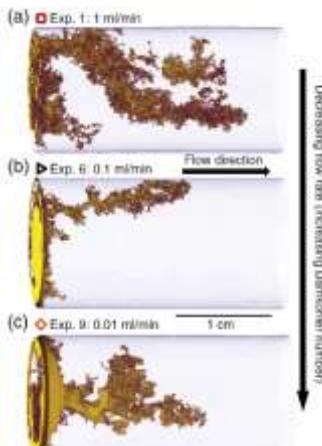
ΔC : concentration.change.between.inlet.and.outlet
 ν : stoichiometric.number.of.the.element.in.K-feldspar
 P : integer.number.of.passes.of.fluid.through.the.core
 S_{ksp} : surface.area.of.K-feldspar
 Q : volumetric.flow.rate.

Involved tasks in DESTRESS

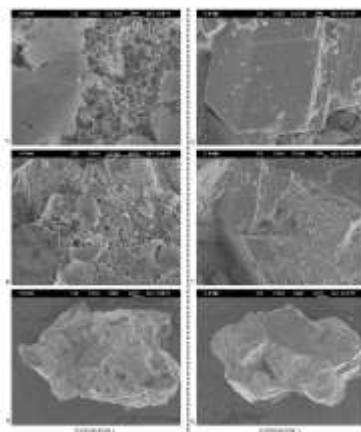


Objectives

Mineral dissolution / precipitation



Luhmann, Kong, et al., 2014

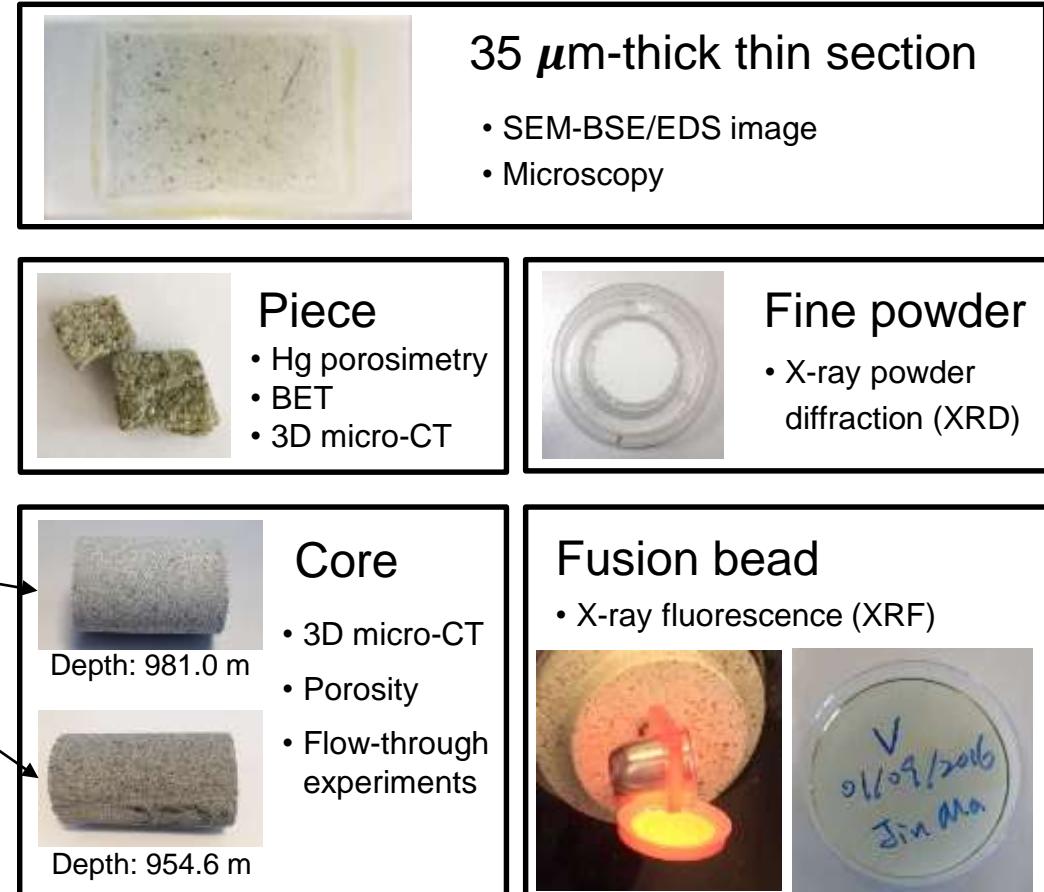


Luhmann, Tutolo, Kong, et al., 2013

Property alterations

- Pore structure
- Pore size distribution
- Porosity
- Permeability
- Flow field
- Accessible surface area
- Reaction rates
- **Reservoir productivity**

Reactive flow-through experiments – Rock samples

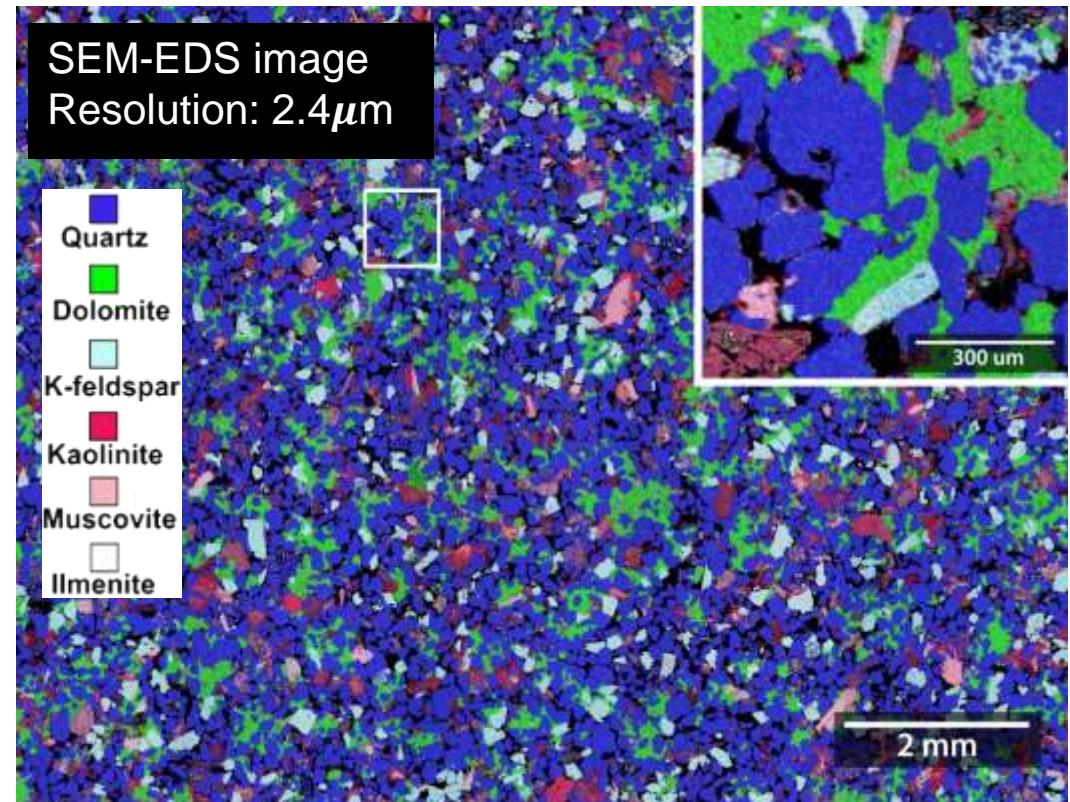


BSE: Backscatter Electron Detector

EDS: Energy Dispersive X-ray Spectroscopy

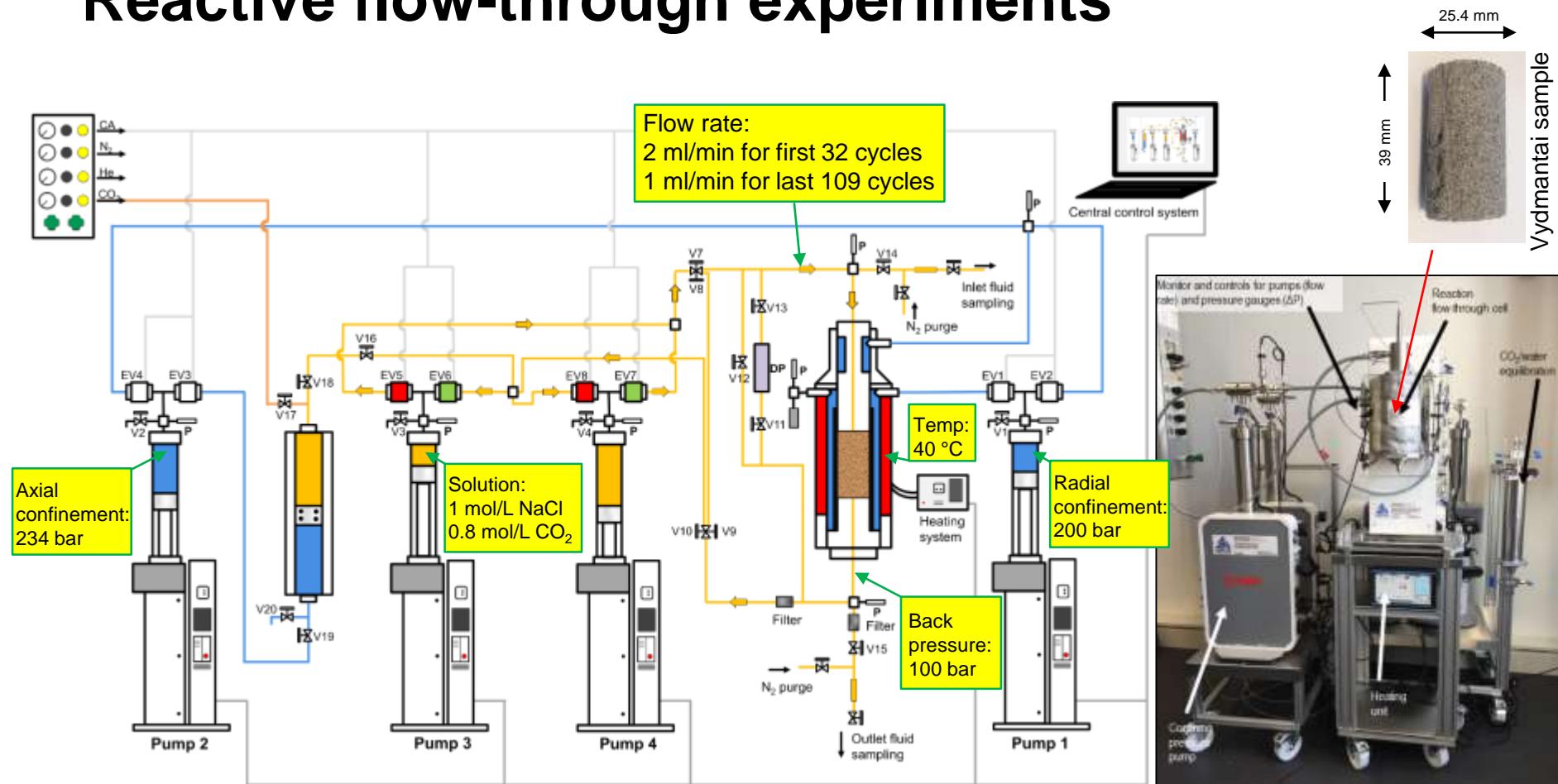
Rock composition

vol. %	Formula
45.53	SiO_2
12.22	$\text{CaMg}_{0.8}\text{Fe}_{0.2}(\text{CO}_3)_2$
9.93	KAISi_3O_8
5.64	$\text{Al}_{1.9}\text{Si}_{2.1}\text{O}_5(\text{OH})_4$
4.76	$\text{K}_{0.5}\text{MgFe}_{1.2}\text{Al}_{0.7}(\text{AlSi}_3\text{O}_8)(\text{OH})_2$
0.27	$\text{Fe}_2\text{Ti}_5\text{O}_{12}$



EDS: Energy Dispersive X-ray Spectroscopy

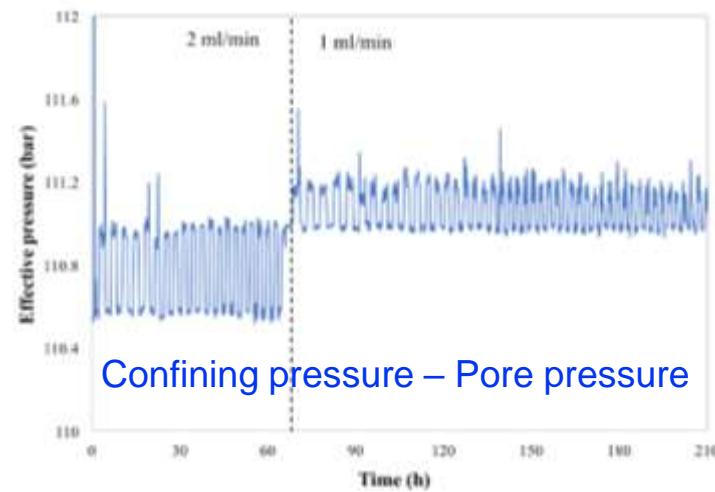
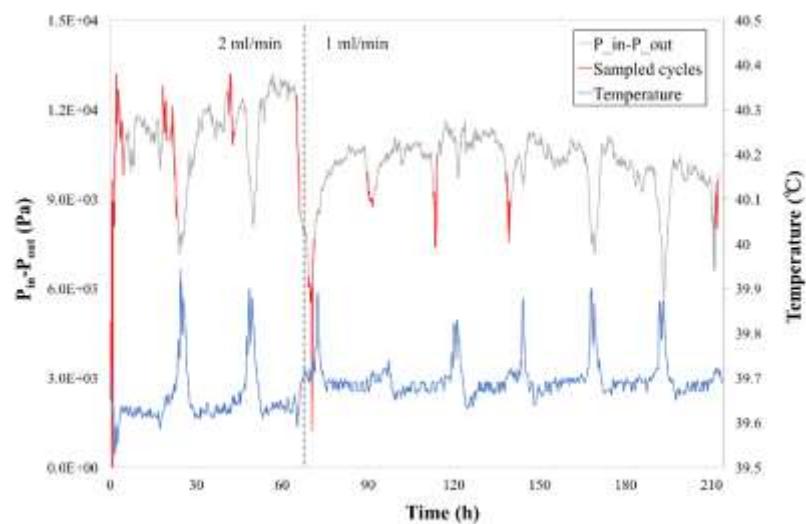
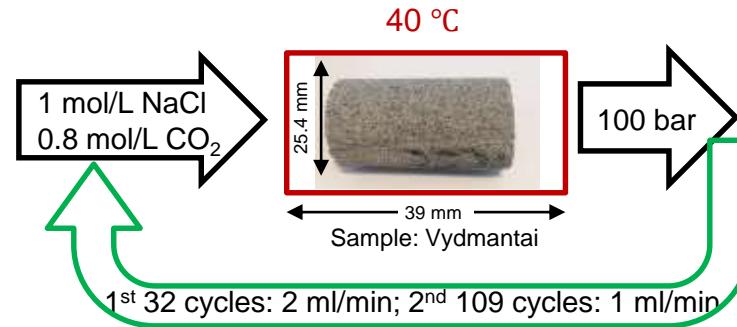
Reactive flow-through experiments



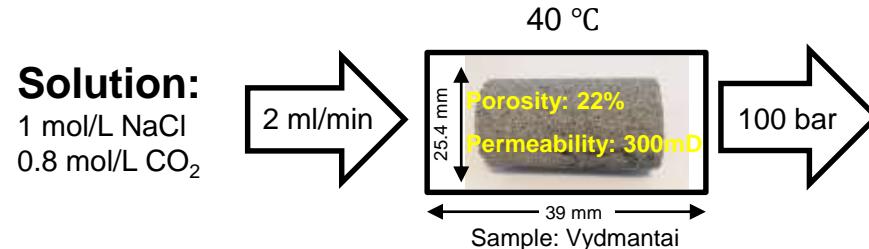
Recycling (RC) experiment: approaching equilibrium

Duration: 10 days
Total cycles: 141
Total recycled volume: 25.2L

Pressure measurements



Numerical simulations with PFLOTRAN



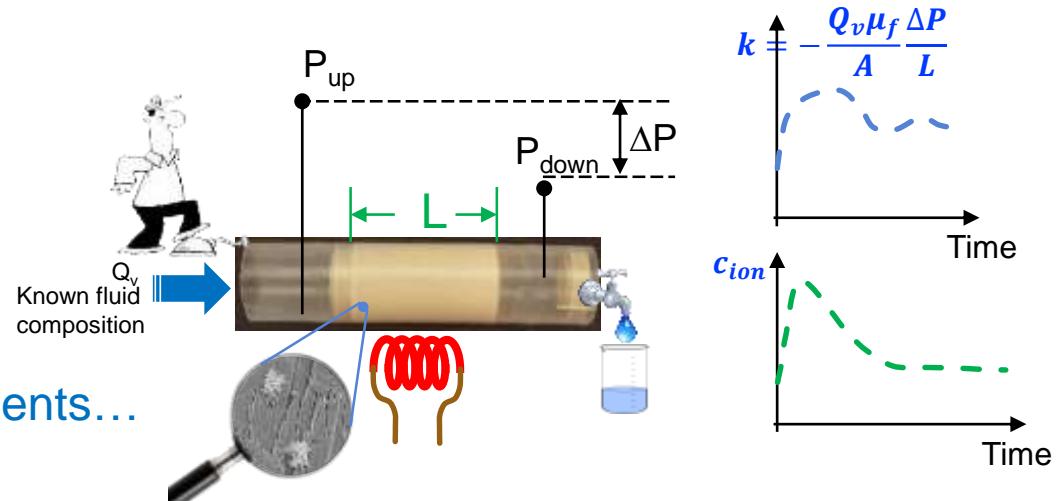
Mineral	Formula	Volume fraction (%)	Reactive surface area (cm ² /cm ³)	Rate constant (mol/cm ² /s) Palandri et al. (2004)
Quartz	SiO ₂	46.0	1521.4	1.0 x 10 ⁻¹⁷
Dolomite	CaMg(CO ₃) ₂	12.0	1328.7	3.0 x 10 ⁻⁸
K-feldspar	KAlSi ₃ O ₈	10.0	180.5	1.0 x 10 ⁻¹²
Muscovite	KAl ₃ Si ₃ O ₁₀ (OH) ₂	5.0	1310.0	1.0 x 10 ⁻¹⁶
Kaolinite	Al ₂ Si ₂ O ₅ (OH) ₄	5.0	26087.2	5.0 x 10 ⁻¹⁶
Ilmenite	FeTiO ₃	0.3	100.2	4.5 x 10 ⁻¹³

Lab measurements

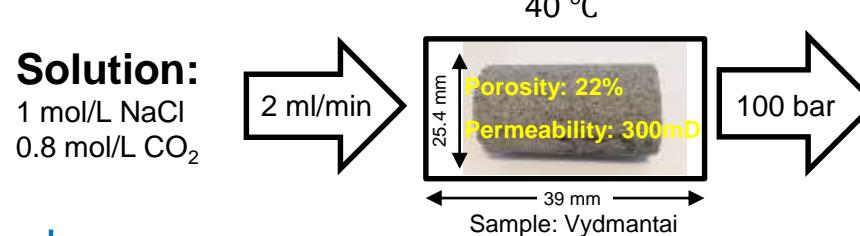
Outlooks

- Capability of flow-through experiments

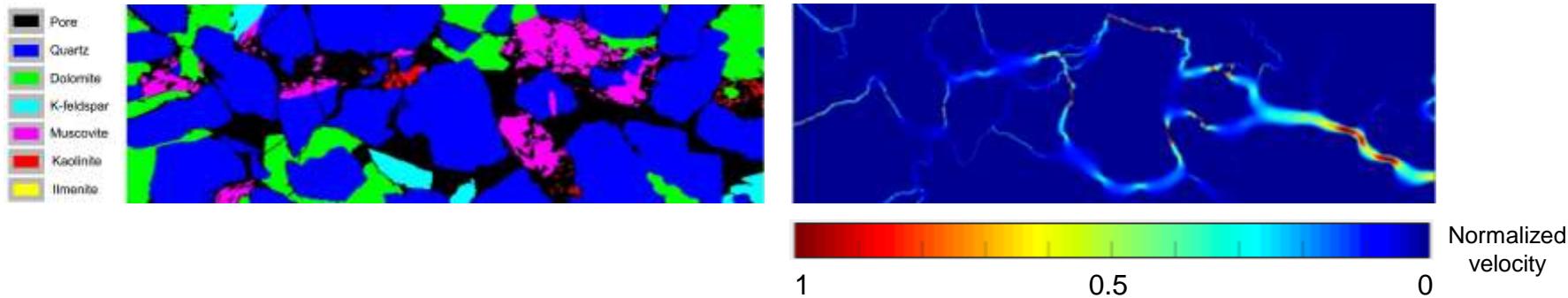
Beyond the experiments...



- Numerical challenges



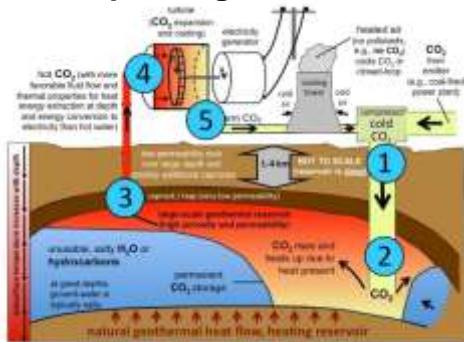
2D and 3D pore-scale



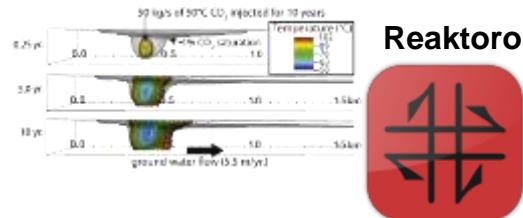
The Geothermal Energy and Geofluids Group

Numerics

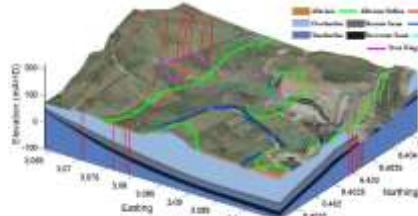
CO₂-plume geothermal



Multiphase-reactive flow

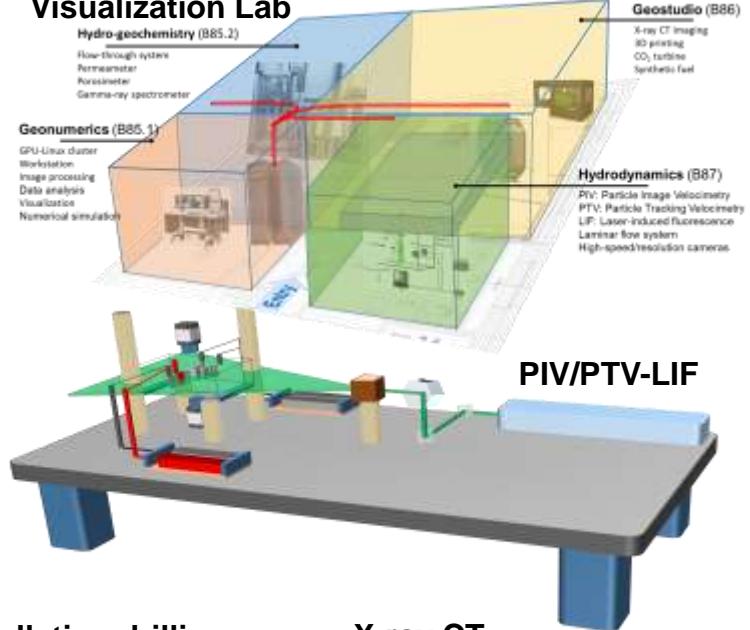


Regional Groundwater Modeling



Laboratory

The Geosystem Reactive Transport (GREAT) Visualization Lab



X-ray CT

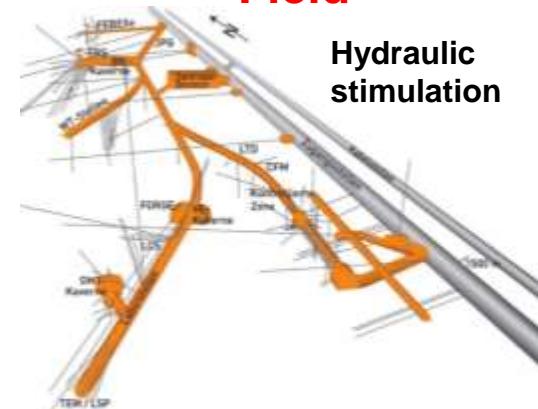


Reactive system

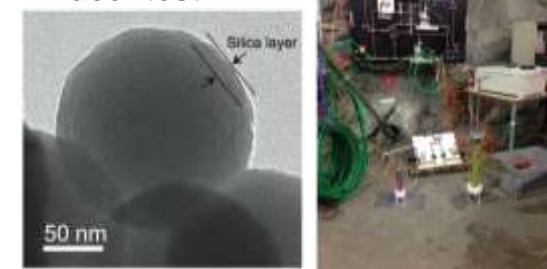


Field

Hydraulic stimulation



Reservoir characterization: Tracer test



Magnetotellurics (MT)

